

**MARK SCHEME for the October/November 2011 question paper  
for the guidance of teachers**

**0652 PHYSICAL SCIENCE**

**0652/21**

Paper 2 (Core Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2011 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.

Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – October/November 2011	0652	21

- 1 (a) balance ; [1]
- (b) burette ; [1]
- (c) thermometer ; [1]
- (d) beaker **OR** burette ; [1]

**[Total: 4]**

- 2 (a) 50 (m/s) ; [1]
- (b) deceleration ;  
constant ; [2]
- (c) use of area under graph,  $S = \frac{1}{2} \times 30 \times 10$  ;  
150 (m) ; [2]  
(calculation  $30 \times 10 = 300 \text{ m} - \text{max } 1$ )
- (d) (i) zero ; [1]
- (ii) mention of frictional force ; [1]
- (e) car **A** ;  
larger gradient ;  
greater acceleration ; [max 2]

**[Total: 9]**

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – October/November 2011	0652	21

3 (a) suitable example of ionic compound e.g. sodium chloride ;  
suitable example of covalent compound e.g. ammonia ; [2]

(b) suitable example for ionic compound ;  
e.g. conduct electricity when molten or in aqueous solution / giant ionic structure  
/ high melting and boiling points / etc.  
  
suitable example for covalent compound ;  
e.g. does not conduct electricity when molten / simple molecular structure  
/ low melting and boiling points / etc. [2]

(c) diagram showing 2 electrons in outer shell ;  
3 shells with 2 electrons in first shell and 8 in middle shell ; [2]

**[Total: 6]**

4 (a) bauxite ; [1]

(b) aluminium too reactive ;  
more reactive than carbon / carbon not reactive enough / will not replace carbon ; [2]

**[Total: 3]**

5 (a) (i) so that the mean temperature of the ice is measured ; [1]

(ii) sample is below room temperature ;  
so absorbs energy from the surroundings ; [2]

(b)  $-2(^{\circ}\text{C})$  ; [1]

(c) temperature remains constant / ice melting ;  
molecules gain potential energy / bonds are broken ; [2]

**[Total: 6]**

<b>Page 4</b>	<b>Mark Scheme: Teachers' version</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>IGCSE – October/November 2011</b>	<b>0652</b>	<b>21</b>

6 (a)

name	formula	mass of 1 mole / g
<b>water</b>	H <sub>2</sub> O	<b>18</b>
hydrogen chloride	<b>HCl</b>	36.5
sodium fluoride	<b>NaF</b>	42
<b>nitrogen</b>	N <sub>2</sub>	<b>28</b>

[4]

(b) Na<sup>+</sup> AND 11 ;

F<sup>-</sup> AND 9 ;

[2]

[Total: 6]

7 (a) (i) 45 ;

[1]

(ii) 60 ;

[1]

(b) (i) (a fast moving) electron ;

[1]

(ii) loses 1 neutron ;  
gains proton ; ('neutron changes to proton' gains 2 marks)

[2]

[Total: 5]

8 (a) suitable advantage, e.g. no pollution, etc. ;  
suitable disadvantage, e.g. needs to be made, etc. ;

[2]

(b)  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  ;;  
(correct formulae – 1 mark and correct balancing – 1 mark)

[2]

(c) lighted splint ;  
pops ;

[2]

(d) (i) ammonia ;

[1]

(ii) Haber/Haber-Bosch ;

[1]

[Total: 8]

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – October/November 2011	0652	21

9 (a) the (vibrating) rubber hits air molecules ;  
causing them to vibrate/forming a sound wave ; [2]  
(no mention of vibration 1 max.)

(b) (i) same frequency (approximately) ; [2]  
smaller amplitude ;

(ii) number of waves (or vibrations) per second ; [2]  
Hz or hertz ;

**[Total: 6]**

10 (a) halogens ; [1]

(b) fluorine/bromine/iodine/astatine ; [1]

(c) correct use of chlorine ; [1]  
e.g. water sterilization/making plastics/etc.

(d) magnesium ; [1]

(e) bubble chlorine into the solution ; [2]  
turns brown / yellow ;

(f) 35 ; [2]  
36 (allow e.c.f. on number in atom, i.e. atom + 1 for a max 1) ;

**[Table: 8]**

11 (a) lamp/bulb ; [1]

(b) (i)  $20 \Omega$  ; [1]

(ii) use of  $I = V / R (= 9 / 20)$  ; [2]  
 $= 0.45 \text{ A}$  ;

(iii) use of  $V = IR (= 0.45 \times 12)$  ; [2]  
 $= 5.4 \text{ V}$  ;

**[Total: 6]**

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – October/November 2011	0652	21

12 (a) alkanes ; [1]

(b) propane ;  
C<sub>3</sub>H<sub>8</sub> ; [2]

(c) contains oxygen ;  
hydrocarbons contain hydrogen and carbon only ; [2]

**[Total: 5]**

13 (a) straight lines between poles ;  
all lines start on one pole and finish on the other, none touch each other ;  
arrows pointing north to south ; [3]

(b) complete circuit ;  
mercury is a conductor ; [2]

(c) the rod will kick ;  
towards/away from the observer ; [2]

(d) kick/move in the opposite direction ; [1]

**[Total: 8]**